**Covalent Compound Properties and Naming**

**Quiz Review Sheet**

1) Why do covalent compounds have lower melting/boiling points than ionic compounds?

2) When two metals bond, a valence electron is transferred from a metal to a nonmetal (as in the case of NaCl, where Na gives a valence electron to Cl). When two nonmetals bond with one another to form a covalent molecule, what happens to the valence electrons during this process? What does the term “molecule” have to do with this?

3) How many electrons are in a covalent bond? Where are these electrons located?

4) Why do ionic compounds generally form when metals bond to nonmetals, while covalent compounds generally form when nonmetals bond to other nonmetals?

5) Name the following: B2H6, NH3, H2O, OF2, Br2

6) Write the formulas of the following compounds: silicon disulfide, silicon tetrachloride, nitrogen trioxide, ammonia, phosphorus